Section-B

(Short Answer)

Note: Answer any EIGHT of the following questions. Each question carries 05 marks.

(Atomic mass : - Fe = 56, O = 16, Na = 23, C = 12)

Q.2 What is the contribution of Jabir Ibn-e-Hayyan in the field of chemistry?

Q.3 The formula of rust is Fe_2O_3. How many moles of Fe are present in 30g of rust?

Q.4 In what way do isotopes of given elements differ from each other?

Q.5 Explain Newland's law of octave. How this law provided the larger scope the classification of the elements?

Q.6 Give the characteristics of Ionic compounds.

Q.7 5.3gm Na_2CO_3 was dissolve din 800gm water. Calculate the molality of solution.

Q.8 Define any TWO of the following terms.

Thermo chemistry _ Exothermic reaction _ Endothermic reaction

Q.9 Defien isotopes: Discuss various isotopes of hydrogen.

Q.10 Give four uses of sulphur.

Q.11 Describe the process of nicle plating.

Q.12 Differentiate between wrought iron and steel.

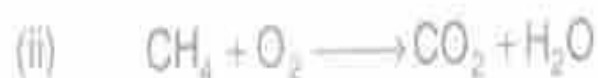
Q.13 Balance the following equations.

(i) $C_2H_2 + H_2 \longrightarrow C_4H_8$

(ii) $CH_4 + O_2 \longrightarrow CO_2 + H_2O$

(iii) $NHO_3 + HCl \longrightarrow NOCl + H_2O + Cl$

(iv) $KClO_3 \longrightarrow KCl + O_2$

(v) $NH_3 + CO_2 \longrightarrow (NH_2)_2CO + H_2O$

Section-C

(Descriptive Answer)

Note: Answer any TWO of the following question. Each question carries 14(7 + 7) marks.

Q.14 (a) How solid is converted into liquid? Explain it.

(b) Comparison between Covalent Bond and Co-ordinate Covalent Bond.

Q.15 (a) What are the sources of halogens? Describe the importance of Cl_2, Br_2 and I_2 in our daily life.

(b) Explain the chain isomersim. Give an example.

Q.16 (a) State and explain Faraday's First law of electrolysis.

(b)Explain the reason why the graphite is good conductor of electricity while dia mond is not, although both are the true allotropes of carbon.